

**Pares redox. Ejercicio nº01, p236. EDEBE.**

Asigna el número de oxidación a cada uno de los elementos en los siguientes compuestos e iones:  $\text{SO}_2$ ,  $\text{K}_2\text{Cr}_2\text{O}_7$ ,  $\text{H}_2\text{CO}_3$ ,  $\text{CH}_2\text{O}$ ,  $\text{MnO}_2$ ,  $\text{Na}_2\text{SO}_4$ ,  $\text{CH}_4$ ,  $\text{BO}_3^{-3}$ ,  $\text{NO}_2^-$ ,  $\text{CrO}_4^{-2}$ ,  $\text{SO}_3^{-2}$ .

+4 -2



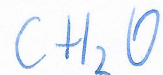
+1 +6 -2



+1 +4 -2



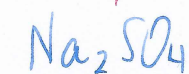
0 +1 -2



+4 -2



+1 +6 -2



-4 +1



+3 -2



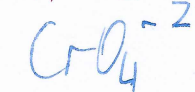
$$3(-2) + X = -3 ; -6 + X = -3 ; X = -3 + 6 = \boxed{+3}$$

+3 -2



$$2(-2) + X = -1 ; -4 + X = -1 ; X = -1 + 4 = \boxed{+3}$$

+6 -2



$$4(-2) + X = -2 ; -8 + X = -2 ; X = -2 + 8 = \boxed{+6}$$

+4 -2 -2



$$3(-2) + X = -2 ; -6 + X = -2 ; X = -2 + 6 = \boxed{+4}$$